Polymers

- A **polymer** is a large molecule with simple repeating units.
- The simple repeating units are called monomers.
- Polymer formulas are described with the formula for the monomer in parentheses with an n as a subscript to indicate some large integer number. The n varies even for a sample of the same polymer.

General polymer formula

Natural and Synthetic Polymers

- Natural polymers include
 - Starches with glucose monomers
 - Proteins with amino acid monomers
- Synthetic polymers are produced by chemists. Examples include, nylon, polyester, polyethylene, poly(vinyl chloride), polypropylene, and polystyrene.

Formation of Ala-Ser-Gly-Cys

peptide bonds (amide functional groups)

Nylon
$$H-N+CH_2 \rightarrow_x N+H + HO \rightarrow C+CH_2 \rightarrow_y C-OH$$

Formation Di-amine Di-carboxylic acid $-H_2O$

repeated many times $\int -H_2O$

$$\begin{pmatrix}
O & O \\
N \leftarrow CH_2 \xrightarrow{x} N - C \leftarrow CH_2 \xrightarrow{y} C \\
H & H \\
Nylon$$

$$n = 40 \text{ to } 110$$

Examples
$$\begin{array}{c} -N + CH_{2} >_{6} N - C + CH_{2} >_{4} C \\ N + CH_{2} >_{6} N - C + CH_{2} >_{4} C \\ N + CH_{2} >_{6} N - C + CH_{2} >_{8} C \\ N >_{8} C + CH_{2} >_{8} C + CH_{2} >_{8} C \\ N >_{8} C + CH_{2} >_{8} C \\ N >_{8} C + CH_{2} >_{8} C \\$$

Condensation Polymers

 Condensation polymers are polymers that are formed by condensation reactions in which two molecules are joined and a small molecule, such as water, is released.

Polyester Formation

Addition Polymers

 Addition polymers are made from molecules that have the following general formula.

$$X$$
 $C = C$
 X
 X
 X

Polyethylene Formation

$$\begin{array}{ccc}
H & H \\
n & C = C \\
H & H \\
ethylene \\
\downarrow & polymerization
\end{array}$$

High- and Low-Density Polyethylene

- If polyethylene is made under conditions that lead to mostly unbranched chains, the chains are able to pack together tightly forming highdensity polyethylene, which is described by the acronym HDPE or sometimes PE-HD.
- If the polyethylene is made in a way that encourages branches, the molecules do not pack together as tightly, forming low-density polyethylene, which is described by the acronym LDPE or sometimes PE-LD.

Poly(vinyl chloride) or PVC

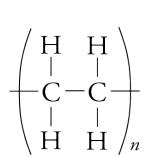
$$\begin{array}{cccc}
H & H \\
n & C = C \\
H & Cl \\
vinyl chloride
\end{array}$$

polymerization

or more simply
$$\begin{pmatrix}
H & H \\
- & - \\
C - C \\
+ & - \\
H & Cl \\
n
\end{pmatrix}$$
 $n = a \text{ very large integer}$

poly(vinyl chloride) or PVC

Addition Polymers



Polyethylene (HDPE or LDPE)

$$\begin{pmatrix}
H & H \\
| & | \\
C - C - \\
| & | \\
H & CH_3/n
\end{pmatrix}$$

Polypropylene (PP)

$$\begin{pmatrix}
H & H \\
| & | \\
C-C \\
| & | \\
H & Cl \\
\end{pmatrix}_{n}$$

Poly(vinyl chloride) (PVC)

Polystyrene (PS)

Recycling Codes

